

## CHAPTER 11

### The p-block elements

Elements in which the last electron enters in the any one of the three p- orbital of their outermost shells – p-block elements

- Gen. electronic configuration of outer shell is  $ns^2np^{1-6}$

The inner core of e-config.may differ which greatly influences their physical & to some extent chemical properties.

- The block of elements in the periodic table consisting of the main groups :
- Group 13 (B to Tl)
- Group 14 (C to Pb)
- Group 15 (N to Bi)
- Group 16 (O to Po)
- Group 17 (F to At)
- Group 18 (He to Rn)

(1) Members at the top and on the right of the *p*-block are nonmetals (C, N, P, O, F, S, Cl, Br, I, At).

(2) Those on the left and at the bottom are metals (Al, Ga, In, Tl, Sn, Pb, Sb, Bi, Po).

(3) Between the two, from the top left to bottom right, lie an ill-defined group of metalloid elements (B, Si, Ge, As, Te)

#### GROUP 13 : The boron group

- Outer Electronic Configuration:- $ns^2np^1$
- group members: boron (B), aluminum (Al), gallium (Ga), indium (In)& thallium (Tl) . All, except boron, are metals.
- Boron show diagonal relationship with Silicon; both are semiconductors metalloids & forms covalent compounds.
- Boron compounds are electron deficient, they are lack of an octet of electrons about the B atom .
- diborane  $B_2H_6$ , is simplest boron hydride

- Structure: three-center two-electron: the H atoms are simultaneously bonded to two B atoms the B-H bridging bond lengths are greater than B-H terminal.
- - Boron oxide is acidic (it reacts readily with water to form boric acid)
- aluminium compounds:aluminium oxide is amphoteric
- aluminum halides, e.g.,  $\text{AlCl}_3$  is dimer, an important catalyst in organic chemistry have anincomplete octet, acts as Lewic acid by accepting lone pairs from Lewic bases, forming adduct
- aluminum hydride, e.g.,  $\text{LiAlH}_4$ , a reducing agent
- Atomic Properties - Electronic Configurations

Element	Symbol	Atomic No.	Electronic Configuration	Abundance in Earth's Crust (in ppm)
Boron	B	5	$[\text{He}]2s^2 2p^1$	8
Aluminium	Al	13	$[\text{Ne}]3s^2 3p^1$	81,300
Galium	Ga	31	$[\text{Ar}]3d^{10}4s^2 4p^1$	15
Indium	In	49	$[\text{Kr}] 4d^{10}5s^2 5p^1$	1
Thallium	Tl	81	$[\text{Xe}] 5d^{10}6s^2 6p^1$	0.3

### . Atomic and ionic radii

- The atomic and ionic radii of group 13 elements are compared to corresponding elements of group 2. From left to right in the period, the magnitude of nuclear charge increases but the electrons are added to, the same shell. These electrons do not screen each other, therefore, the electrons experience greater nuclear charge.
- In other words, effective nuclear charge increases and thus, size decreases. Therefore, the elements of this group have smaller size than the corresponding elements of second group.
- On moving down the group both atomic and ionic radii are expected to increase due to the addition of new shells. However, the observed atomic radius of Al (143 pm) is slightly more than that of Ga (135 pm).

### Ionization energies

The first ionization energies of group 13 elements are less than the corresponding members of the alkaline earths.

The sharp decrease in I.E. from B to Al is due to increase in size. In case of Ga, there are ten d-electrons in its inner electronic configuration.

The very high value of 3<sup>rd</sup> I. E. of thallium indicates that +3 O.N. state is not stable, rather +1 is more stable for thallium .

Electropositive (or metallic) character

the elements of group 13 are less electropositive as compared to elements of group 2. On moving down the group the electropositive (metallic) character increases because ionization energy decreases. For e.g., Boron is a non-metal while the other elements are typical metals.

Oxidation states

The common oxidation states of group 13 elements are +3 and +1 .The stability of the + 1 oxidation state increases in the sequence Al < Ga < In < Tl, Due to Inert pair effect.

Element	B	Al	Ga	In	Tl
Oxidation state	+3	+3	+3, +1	+3, +1	+3, +1

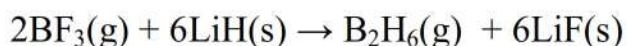
Chemical reactivity of Gr.13 Elements

All elements in their compounds exhibit the oxidation state of + 3 and +1.

Hydrides

- None of the group 13 elements reacts directly with hydrogen. However, a no. of hydrides of these elements have been prepared by indirect methods. The boron hydrides are called boranes & classified in two series: (a)  $B_nH_{n+4}$  called nido-boranes (b)  $B_nH_{n+6}$  called arachno-boranes

- INDUSTRIAL PREPERATION :-



- Laboratory method:

- By the reaction of iodine with sodium borohydride in a high boiling solvent.

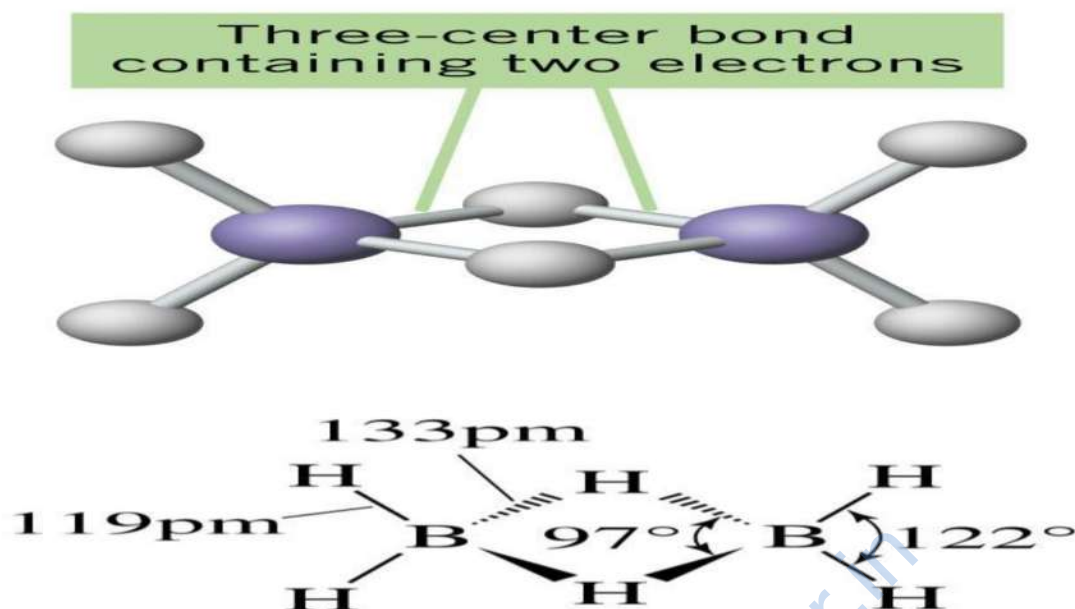


- By reduction of  $BCl_3$  with  $LiAlH_4$





## Structure of Diborane, B<sub>2</sub>H<sub>6</sub>



Some important characteristics of boranes:

- Lower boranes are colourless gases while higher boranes are volatile liquids or solids.
- They undergo spontaneous combustion in air due to strong affinity of boron for oxygen.

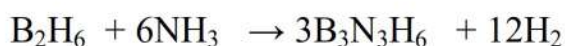


- Boranes react with alkali metal hydrides in diethyl ether to form borohydride complexes.



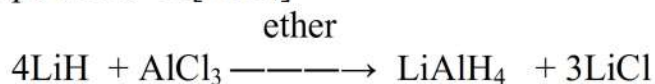
Metal borohydride

- (iv) Diborane reacts with ammonia to give borazine at 450 K.



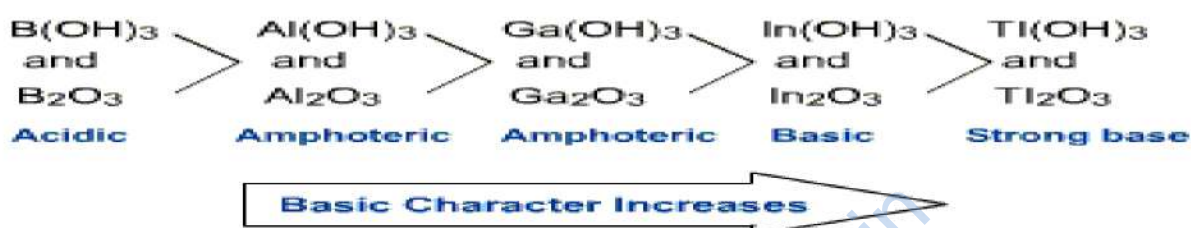
- Borazine has a cyclic structure similar to benzene and thus is called inorganic benzene
- The other elements of this group form only a few stable hydrides. The thermal stability decreases as we move down the group.

- $\text{AlH}_3$  is a colourless solid polymerized via Al - H - Al bridging units. These hydrides are weak Lewis acids and readily form adducts with strong Lewis base (B:) to give compounds of the type  $\text{MH}_3$  (M = Al or Ga). They also form complex-tetrahydrido anions,  $[\text{MH}_4]^-$ . The most important tetrahydrido compound is  $\text{Li}[\text{AlH}_4]$



## OXIDES & HYDROXIDES

- $\text{M}_2\text{O}_3$  &  $\text{M}(\text{OH})_3$



## HALIDES: Structure of boron trihalides



## Dimeric structure of aluminium chloride

- Boron halides do not form dimers because the size of boron is so small that it is unable to coordinate four large-sized halide ions.

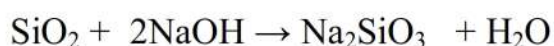
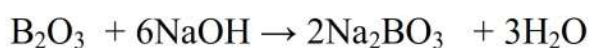


- Anomalous properties of boron

1. Boron is a non-metal & bad conductor of electricity whereas aluminium is a metal & good conductor. B is hard but Al is a soft metal.
2. Boron exists in two forms-crystalline and amorphous. But Al does not exist in different forms.

3. The melting and boiling point of boron are much higher than that of Al .
4. Boron forms only covalent compounds whereas Al forms even some ionic compounds.
5. The hydroxides and oxides of boron are acidic in nature whereas those of aluminium are amphoteric.
6. The trihalides of boron exist as monomers. On the other hand, aluminium halides exist as dimers .
7. The hydrides of boron are quite stable while those of aluminium are unstable

- Boron and silicon exhibit the typical properties of non-metals. These do not form cations. Both exist in amorphous as well as crystalline forms.
- Boron oxide ( $B_2O_3$ ) and silica ( $SiO_2$ ) both are acidic and dissolve in alkali solutions to form borates and silicates respectively.



- The chlorides of both B and Si get hydrolyzed by water to boric acid and silicic acid respectively.



- The hydrides of Boron and Silicon are quite stable. Numerous volatile hydrides are also known which catch fire on exposure to air and are easily hydrolyzed.
- Both elements are semiconductors.

### Behavior in Aqueous Solutions

1 Al, Ga, In and Tl exhibit a well-defined aqueous chemistry in their tripositive states. Species like  $[M(OH)_4]^-$ ,  $[M(H_2O)_2(OH)_4]^-$ ,  $[M(OH)_6]^{3+}$  for  $M = Al, Ga, In$ , exist in aqueous solution.

2. Al, Ga, In and Tl ions exist as octahedral aqua ions,  $[M(OH)_6]^{3+}$  in aqueous solution and many salts like halides, sulphates, nitrates and perchlorates exist as hydrates.

3. Aluminiumsulphate forms double salts - called alum, having the general formula



### USES OF BORON & ALUMINIUM



- Aluminium is used extensively in industry and everyday life. It forms many useful alloys with Cu, Mn, Mg, Si and Zn. Hence, aluminium and its alloys find use in packaging, utensil making, construction, aerospace and other transportation industries. It is used as a conductor for transmission of electricity. Aluminium is also used in the aluminothermite process for production of chromium and manganese from their ores.

### Group 14 Elements:-The Carbon Family

Group 14 includes carbon (C), silicon (Si), germanium (Ge), tin (Sn) and lead (Pb).

General electronic configuration of carbon family is  $ns^2np^2$ .

Covalent radius:-Covalent radius expected to increase from C to Si,

From Si to Pb small increase is found.

Ionization Enthalpy:-The first ionization enthalpies of group 14 elements are higher than those of the corresponding group 13 elements.

Electronegativity:-Group 14 elements are smaller in size as compared to group 13 elements that's why this group elements are slightly more electronegative than group 13

Chemical properties:-

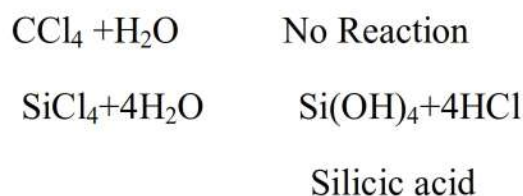
Carbon and silicon mostly show +4 oxidation state. Germanium forms stable compounds in +4 state and only few compounds in +2 state.

Tin forms compounds in both oxidation states. Lead compounds in +2 state are stable and in +4 state are strong oxidizing agents.

Exception:- $Pb_4$  and  $SnF_4$  are ionic in nature.

Except  $CCl_4$  other tetrachlorides are easily hydrolysed by water.

Since carbon does not have d-orbitals and hence cannot expand its coordination number beyond 4



Allotropes of Carbon:-The three types of allotropes are –

1-Diamond

2-Graphite

3-Fullerence

Diamond:-In diamond each carbon atom undergoes  $SP^3$  hybridisation.

Each carbon is tetrahedrally linked to four other carbon atoms.

Graphite:-In graphite, carbon is  $SP^2$ -hybridized. Graphite has a two-dimensional sheet like structure consisting of a number of hexagonal rings fused together.

Graphite conducts electricity along the sheet. It is very soft and slippery.

Fullerene was discovered collectively by three scientists namely R.E Smalley, R.F Curl and H.W Kroto.

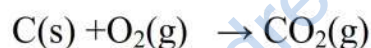
SOME Important Compounds Of Carbon and Silicon

Carbon monoxide:-It is prepared by direct oxidation of C in limited supply of oxygen.



Commercially it is prepared by the passage of steam over hot coke.

Carbon Dioxide:-It is prepared by complete combustion of carbon and carbon fuels in excess of air.



Laboratory method:-

In laboratory it is prepared by the treatment of dil HCl on  $CaCO_3$ .



Silicon dioxide:-Silicon dioxide is a COVALENT THREE DIMENSIONAL NETWORK SOLID.

Each silicon atom is covalently bonded in a tetrahedral manner to four oxygen atoms.

Silicones:-Silicones are the synthetic organo-silicon polymers having general formulae  $(R_2SiO)_n$  in which R = alkyl (methyl, ethyl or phenyl).

Silicates:-Silicates exist in nature in the form of feldspar, zeolites, mica and asbestos etc.

The basic structure of silicates is  $SiO_4^{4-}$ .



Zeolites:-Zeolites is aalumino-silicate of metal. Metal cations participating in formationof Zeolite are use usually  $\text{Na}^+$ ,  $\text{K}^+$ , or  $\text{Ca}^{2+}$ .

Zeolites are used to remove permanent hardness of water.

### ONE MARK QUESTIONS

1. Why is boron used in nuclear reactions?

Ans:-Because Boron can absorb neutrons.

2. By giving a balanced equation show how  $\text{B}(\text{OH})_3$  behaves as an acid in water.



3.Name the element of group 14 which exhibits maximum tendency for catenation?

Ans:-Carbon

4. What is the basic building unit of all silicates?

Ans:- $\text{SiO}_4^{4-}$  is the basic unit of all silicates.

5. What happens when  $\text{NaBH}_4$  reacts with iodine?



6. What happens when boric acid is heated



7. What is producer gas?

Ans:-Producer gas is a mixture of CO and  $\text{N}_2$  in the ratio of 2:1.

8. Write the state of hybridization of 'B' in  $\text{BF}_3$ .

ANS:-Hybridisation of 'B' in  $\text{BF}_3$  is  $\text{Sp}^2$ .

9.Mention the state of hybridization in B in  $\text{BH}_4^-$ .

Ans:- $\text{Sp}^3$ .

10. Which oxide of carbon is regarded as anhydride of carbonic acid.

Ans:- $\text{CO}_2$  is regarded as a hydride of carbonic acid .



## TWO MARKS QUESTIONS

1. Give the chemical reaction as an evidence for each of the following observations.

(i) Tin (II) is a reducing agent whereas lead (II) is not.

(ii) Gallium (I) undergoes disproportionation reaction.

Ans:- (i) Due to inert pair effect  $Pb^{2+}$  is more stable than  $Pb^{4+}$ . Whereas  $Sn^{4+}$  is more stable than  $Sn^{2+}$ .



This is because  $Ga^{3+}$  is more stable than  $Ga^+$ .

2. What happens when

(i) Quick lime is heated with coke?

(ii) Carbon monoxide reacts with  $Cl_2$



3. Give reason

(i) C and Si are always tetravalent but Ge, Sn, Pb show divalency.

(ii) Gallium has higher ionization enthalpy than Al. Explain.

Ans:- (i) Ge, Sn, Pb show divalency due to inert pair effect,  $Pb^{2+}$  is more stable than  $Pb^{4+}$ .

(ii) Due to poor shielding effect of d-electrons in Ga effective nuclear charge increases as compared to Al thus the I.E is higher than Al.

4. Give reason why boron and aluminium tend to form covalent compounds.

Ans:- Sum of three ionization of both the elements are very high. Thus they have no tendency to lose electrons to form ionic compounds. Instead they form covalent compounds.

5. If B-Cl bond has a dipole moment, explain why  $BCl_3$  molecule has zero dipole moment.

Ans:- B-Cl bond has dipole moment because of polarity. In  $\text{BCl}_3$  since the molecule is symmetrical thus the polarities cancel out.

6. Suggest a reason as to why CO is poisonous.

Ans:- CO reacts with haemoglobin to form carboxy-haemoglobin which can destroy the oxygen carrying capacity of haemoglobin and the man dies of suffocation.

7. What do you understand by-

(a) Inert pair effect:- The pair of electron in the valence shell does not take part in bond formation it is called inert pair effect.

(b) Allotropy:- It is the property of the element by which an element can exist in two forms which have same chemical properties but different physical properties due to their structures.

8. How is excessive content of  $\text{CO}_2$  responsible for global warming?

Ans:- Excess of  $\text{CO}_2$  absorbs heat radiated by the earth. Some of it is dissipated into the atmosphere while the remaining part is radiated back to the earth. Temperature of the earth increases.

9. Describe two similarities and two dissimilarities between B and Al.

Ans:- Similarities:-

- (i) Both have same number of valence electrons.
- (ii) Both have similar electronic configuration.

Dissimilarities:-

- (i) B is a non-metal where Al is a metal
- (ii) B forms acidic oxide whereas Al forms amphoteric oxides.

10. What are fullerenes? How were they prepared?

Ans:- Fullerenes are the allotropes of carbon. Its structure is like a soccer ball.

They are prepared by heating graphite in electric arc in presence of inert gases such as helium or argon.

### THREE MARKS QUESTIONS

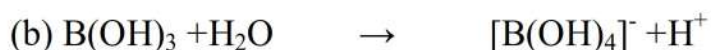
1. What happens when

- (a) Borax is heated strongly



(b) Boric acid is added to water

(c) Aluminium is treated with dilute NaOH



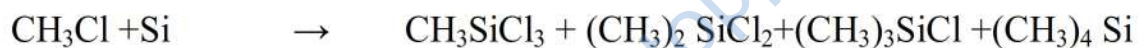
2. Explain the following reactions.

(a) Silicon is heated with methyl chloride at high temperature in the presence of copper.

(b) Silicon dioxide is treated with hydrogen fluoride.

(c) CO is heated with ZnO.

Ans:- (a) A mixture of mono-, di- and trimethylchlorosilanes along with a small amount of tetramethylsilane is formed.



(b) The initially formed silicon tetrafluoride dissolves in HF to form hydrofluorosilicic acid



(c) ZnO is reduced to zinc metal



3. Give reasons:-

(a) Diamond is used as an abrasive.

(b) Aluminium alloys are used to make aircraft body.

(c) Aluminium utensils should not be kept in water overnight.

Ans:- (a) Diamond is used as an abrasive because it is an extremely hard substance.

(b) Alloys of aluminium like duralium is used to make aircraft body due to some of its properties.

(c) Generally aluminium metal does not react with water quickly but when it is kept overnight. It reacts slowly with water in presence of air.

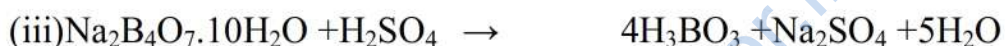


4. A certain salt X, gives the following results.

(i) Its aqueous solution is alkaline to litmus.

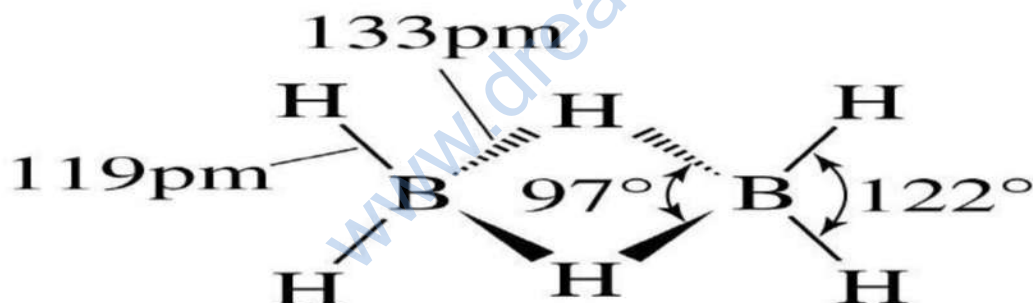
(ii) It swells up to a glassy material Y on strong heating.

(iii) when conc.  $\text{H}_2\text{SO}_4$  is added to a hot solution of X, white crystal of an acid Z separates out.



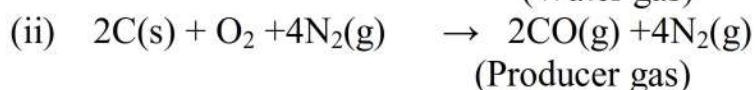
5. draw structure of diborane .

Ans.



### FIVE MARKS QUESTIONS

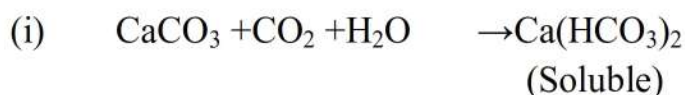
1 Explain the formation of (i) Water gas (ii) Producer gas. Give their uses. What happens when  $\text{CO}_2$  is passed through lime water (i) for short duration (ii) for long duration.



Water gas and Producer gas are used as fuel.



(White ppt.)



- 2 (a) Why do Boron halides form addition compound with  $\text{NH}_3$  ?  
(b) Assign appropriate reason for each of the following observations :-

- (i) Anhydrous  $\text{AlCl}_3$  is used as a catalyst in many organic reactions.  
(ii) No form of elemental silicon is comparable to graphite.

Ans:- (a) It is because  $\text{BX}_3$  is electron deficient whereas  $\text{NH}_3$  is electron rich.

(b) (i) It is Lewis acid.

(ii) It cannot form  $p\pi - p\pi$  bond due to large size.

3. (i) Give reason for the following observations:-  
(a) The tendency for catenation decreases down the group in Group 14.  
(b) The decreasing stability of +3 oxidation state with increasing atomic number in group 13.  
(c)  $\text{PbO}_2$  is a stronger oxidizing agent than  $\text{SnO}_2$ .  
(d) Molten aluminium bromide is a poor conductor of electricity.

Ans:- (i)(a) It is due to decrease in bond dissociation energy which is due to increase in atomic size.

$\text{C-C} > \text{Si-Si} > \text{Ge-Ge} > \text{Sn-Sn} > \text{Pb-Pb}$ .

(b) It is due to inert pair effect.

(c)  $\text{PbO}_2$  is stronger oxidizing agent than  $\text{SnO}_2$  because  $\text{Pb}^{2+}$  is more stable than  $\text{Pb}^{4+}$  whereas  $\text{Sn}^{4+}$  is more stable than  $\text{Sn}^{2+}$ .

(d) Molten  $\text{AlBr}_3$  is poor conductor of electricity because it is covalent compound.