

# 13

## Nuclei

### Multiple Choice Questions (MCQs)

**Q. 1** Suppose we consider a large number of containers each containing initially 10000 atoms of a radioactive material with a half life of 1 yr. After 1 yr,

- (a) all the containers will have 5000 atoms of the material
- (b) all the containers will contain the same number of atoms of the material but that number will only be approximately 5000
- (c) the containers will in general have different numbers of the atoms of the material but their average will be close to 5000
- (d) none of the containers can have more than 5000 atoms

**K Thinking Process**

$$T_{1/2} = \frac{\ln 2}{\lambda}, \lambda \rightarrow \text{decay constant.}$$

**Ans. (c)** Radioactivity is a process due to which a radioactive material spontaneously decays. In half-life ( $t = 1\text{yr}$ ) of the material on the average half the number of atoms will decay.

Therefore, the containers will in general have different number of atoms of the material, but their average will be approx 5000.

**Q. 2** The gravitational force between a H-atom and another particle of mass  $m$  will be given by Newton's law

$$F = G \frac{M \cdot m}{r^2}, \text{ where } r \text{ is in km and}$$

- (a)  $M = m_{\text{proton}} + m_{\text{electron}}$
- (b)  $M = m_{\text{proton}} + m_{\text{electron}} - \frac{B}{c^2}$  ( $B = 13.6 \text{ eV}$ ).
- (c)  $M$  is not relate to the mass of the hydrogen atom.
- (d)  $M = m_{\text{proton}} + m_{\text{electron}} - \frac{|V|}{c^2}$  ( $|V| =$  magnitude of the potential energy of electron in the H-atom).

**Ans (b)** Given,  $F = \frac{GMm}{r^2}$

$M$  = effective mass of hydrogen atom

$$= \text{mass of electron} + \text{mass of proton} - \frac{B^2}{C}$$

where  $B$  is BE of hydrogen atom = 13.6 eV.

**Q. 3** When a nucleus in an atom undergoes a radioactive decay, the electronic energy levels of the atom

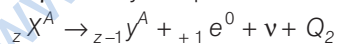
- (a) do not change for any type of radioactivity
- (b) change for  $\alpha$  and  $\beta$ -radioactivity but not for  $\gamma$ -radioactivity
- (c) change for  $\alpha$ -radioactivity but not for others
- (d) change for  $\beta$ -radioactivity but not for others

**Ans. (b)**  $\alpha$ - $\beta$  particle carries one unit of negative charge, an  $\alpha$ -particle carries 2 units of positive charge and  $\gamma$  (particle) carries no charge, therefore electronic energy levels of the atom changes for  $\alpha$  and  $\beta$  decay, but not for  $\gamma$ -decay.

**Q. 4**  $M_x$  and  $M_y$  denote the atomic masses of the parent and the daughter nuclei respectively in radioactive decay. The  $Q$ -value for a  $\beta^-$  decay is  $Q_1$  and that for a  $\beta^+$  decay is  $Q_2$ . If  $m_e$  denotes the mass of an electron, then which of the following statements is correct?

- (a)  $Q_1 = (M_x - M_y) c^2$  and  $Q_2 = [M_x - M_y - 2m_e] c^2$
- (b)  $Q_1 = (M_x - M_y) c^2$  and  $Q_2 = (M_x - M_y) c^2$
- (c)  $Q_1 = (M_x - M_y - 2m_e) c^2$  and  $Q_2 = (M_x - M_y + 2m_e) c^2$
- (d)  $Q_1 = (M_x - M_y + 2m_e) c^2$  and  $Q_2 = (M_x - M_y + 2m_e) c^2$

**Ans. (a)** Let the nucleus is  ${}_z X^A$ .  $\beta^+$  decay is represented as



$$\begin{aligned} \therefore Q_2 &= [m_n({}_z X^A) - m_n({}_{z-1} Y^A) - m_e] c^2 \\ &= [m_n({}_z X^A) + z m_e - m_n({}_{z-1} Y^A) - (z-1) m_e - 2m_e] c^2 \\ &= [m({}_z X^A) - m({}_{z-1} Y^A) - 2m_e] c^2 \\ &= (M_x - M_y - 2m_e) c^2 \end{aligned}$$

$\beta^-$  decay is represented as

$$\begin{aligned} &= {}_z X^A \rightarrow {}_{z+1} A^Y + {}_{-1} e^0 + \bar{\nu} + \alpha_1 \\ \alpha_1 &= [m_n({}_z X^A) - m_n({}_{z+1} Y^A) - m_e] c^2 \\ &= [m_n({}_z X^A) + z m_e - m_n({}_{z+1} Y^A) - (z+1) m_e] c^2 \\ &= [m({}_z X^A) - m({}_{z+1} Y^A)] c^2 \\ &= (M_x - M_y) c^2 \end{aligned}$$

**Q. 5** Tritium is an isotope of hydrogen whose nucleus triton contains 2 neutrons and 1 proton. Free neutrons decay into  $p + \bar{e} + \bar{n}$ . If one of the neutrons in Triton decays, it would transform into  $\text{He}^3$  nucleus. This does not happen. This is because

- (a) Triton energy is than that of a  $\text{He}^3$  nucleus
- (b) The electron created in the beta decay process cannot remain in the nucleus
- (c) both the neutrons in triton have to decay simultaneously resulting in a nucleus with 3 protons, which is not a  $\text{He}^3$  nucleus.
- (d) free neutrons decay due to external perturbations which is absent in triton nucleus

**K Thinking Process**

*Isotopes of an element are having same atomic numbers and different mass numbers.*

**Ans. (a)** Tritium ( ${}_1\text{H}^3$ ) contains 1 proton and 2 neutrons. A neutron decays as  $n \longrightarrow p + \bar{e} + \bar{\nu}$ , the nucleus may have 2 protons and one neutron, i.e., tritium will transform into  $2\text{He}^3$  (2 protons and 1 neutron).

Triton energy is less than that of  $2\text{He}^3$  nucleus, i.e., transformation is not allowed energetically.

**Q. 6** Heavy stable nuclei have more neutrons than protons. This is because of the fact that

- (a) neutrons are heavier than protons
- (b) electrostatic force between protons are repulsive
- (c) neutrons decay into protons through beta decay
- (d) nuclear forces between neutrons are weaker than that between protons

**Ans. (b)** Stable heavy nuclei have more neutrons than protons. This is because electrostatic force between protons is repulsive, which may reduce stability.

**Q. 7** In a nuclear reactor, moderators slow down the neutrons which come out in a fission process. The moderator used have light nuclei. Heavy nuclei will not serve the purpose, because

- (a) they will break up
- (b) elastic collision of neutrons with heavy nuclei will not slow them down
- (c) the net weight of the reactor would be unbearably high
- (d) substances with heavy nuclei do not occur in liquid or gaseous state at room temperature

**K Thinking Process**

*When there is an elastic collision between two bodies of same mass their velocities are exchanged.*

**Ans. (b)** According to the question, the moderator used have light nuclei (like proton). When protons undergo perfectly elastic collision with the neutron emitted their velocities are exchanged, i.e., neutrons come to rest and protons move with the velocity of neutrons. Heavy nuclei will not serve the purpose because elastic collisions of neutrons with heavy nuclei will not slow them down.

## Multiple Choice Questions (More Than One Options)

**Q. 8** Fusion processes, like combining two deuterons to form a He nucleus are impossible at ordinary temperatures and pressure. The reasons for this can be traced to the fact

- (a) nuclear forces have short range
- (b) nuclei are positively charged
- (c) the original nuclei must be completely ionized before fusion can take place
- (d) the original nuclei must first break up before combining with each other

**Ans. (a, b)**

Fusion processes are impossible at ordinary temperatures and pressures. The reason is nuclei are positively charged and nuclear forces are short range strongest forces.

**Q. 9** Samples of two radioactive nuclides  $A$  and  $B$  are taken  $\lambda_A$  and  $\lambda_B$  are the disintegration constants of  $A$  and  $B$  respectively. In which of the following cases, the two samples can simultaneously have the same decay rate at any time?

- (a) Initial rate of decay of  $A$  is twice the initial rate of decay of  $B$  and  $\lambda_A = \lambda_B$
- (b) Initial rate of decay of  $A$  is twice the initial rate of decay of  $B$  and  $\lambda_A > \lambda_B$
- (c) Initial rate of decay of  $B$  is twice the initial rate of decay of  $A$  and  $\lambda_A > \lambda_B$
- (d) Initial rate of decay of  $B$  is same as the rate of decay of  $A$  at  $t = 2h$  and  $\lambda_B < \lambda_A$

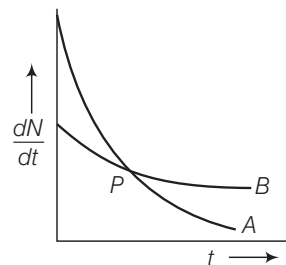
**Ans. (b, d)**

The two samples of the two radioactive nuclides  $A$  and  $B$  can simultaneously have the same decay rate at any time if initial rate of decay of  $A$  is twice the initial rate of decay of  $B$  and  $\lambda_A > \lambda_B$ . Also, when initial rate of decay of  $B$  is same as rate of decay of  $A$  at  $t = 2h$  and  $\lambda_B < \lambda_A$ .

**Q. 10** The variation of decay rate of two radioactive samples  $A$  and  $B$  with time is shown in figure.

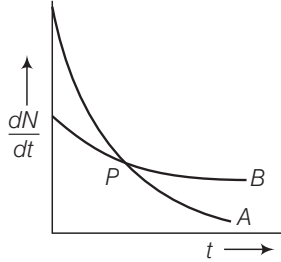
Which of the following statements are true?

- (a) Decay constant of  $A$  is greater than that of  $B$ , hence  $A$  always decays faster than  $B$
- (b) Decay constant of  $B$  is greater than that of  $A$  but its decay rate is always smaller than that of  $A$
- (c) Decay constant of  $A$  is greater than that of  $B$  but it does not always decay faster than  $B$
- (d) Decay constant of  $B$  is smaller than that of  $A$  but still its decay rate becomes equal to that of  $A$  at a later instant



**Ans. (c, d)**

From the given figure, it is clear that slope of curve A is greater than that of curve B. So rate of decay is faster for A than that of B.



We know that  $\left(\frac{dN}{dt}\right) \propto \lambda$ , at any instant of time hence we can say that  $\lambda_A > \lambda_B$ . At point P shown in the diagram the two curves intersect. Hence at point P, rate of decay for both A and B is the same.

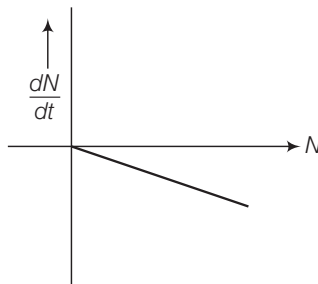
## Very Short Answer Type Questions

**Q. 11**  $\text{He}_2^3$  and  $\text{He}_1^3$  nuclei have the same mass number. Do they have the same binding energy?

**Ans.** Nuclei  $\text{He}_2^3$  and  $\text{He}_1^3$  have the same mass number.  $\text{He}_2^3$  has two protons and one neutron.  $\text{He}_1^3$  has one proton and two neutrons. The repulsive force between protons is missing in  ${}_1\text{He}^3$  so the binding energy of  ${}_1\text{He}^3$  is greater than that of  ${}_2\text{He}^3$ .

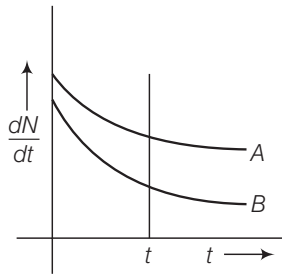
**Q. 12** Draw a graph showing the variation of decay rate with number of active nuclei.

**Ans.** We know that, rate of decay =  $\frac{-dN}{dt} = \lambda N$



where decay constant ( $\lambda$ ) is constant for a given radioactive material. Therefore, graph between  $N$  and  $\frac{dN}{dt}$  is a straight line as shown in the diagram.

**Q. 13** Which sample A or B shown in figure has shorter mean-life?



**Ans.** From the given figure, we can say that

$$\text{at } t = 0, \left(\frac{dN}{dt}\right)_A = \left(\frac{dN}{dt}\right)_B$$

$$\Rightarrow (N_0)_A = (N_0)_B$$

Considering any instant  $t$  by drawing a line perpendicular to time axis, we find that

$$\left(\frac{dN}{dt}\right)_A > \left(\frac{dN}{dt}\right)_B$$

$\Rightarrow$

$$\lambda_A N_A > \lambda_B N_B$$

$\therefore$

$$N_A > N_B$$

(rate of decay of B is slower)

$\therefore$

$$\lambda_B > \lambda_A$$

$\Rightarrow$

$$\tau_A > \tau_B$$

$$\left[ \because \text{Average life } \tau = \frac{1}{\lambda} \right]$$

**Q. 14** Which one of the following cannot emit radiation and why? Excited nucleus, excited electron.

**Ans.** Excited electron cannot emit radiation because energy of electronic energy levels is in the range of eV and not MeV (mega electron volt).

$\gamma$ -radiations have energy of the order of MeV.

**Q. 15** In pair annihilation, an electron and a positron destroy each other to produce gamma radiations. How is the momentum conserved?

**Ans.** In pair annihilation, an electron and a positron destroy each other to produce  $2\gamma$  photons which move in opposite directions to conserve linear momentum. The annihilation is shown below  ${}_0e^{-1} + {}_0e^{+1} \rightarrow 2\gamma$  ray photons.

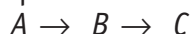
## Short Answer Type Questions

**Q. 16** Why do stable nuclei never have more protons than neutrons?

**Ans.** Because protons are positively charged and repel one another electrically.

This repulsion becomes so great in nuclei with more than 10 protons or so, that an excess of neutrons which produce only attractive forces, is required for stability.

**Q. 17** Consider a radioactive nucleus  $A$  which decays to a stable nucleus  $C$  through the following sequence

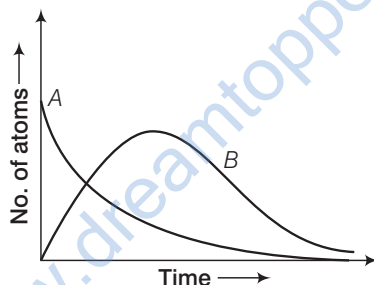


Here  $B$  is an intermediate nuclei which is also radioactive. Considering that there are  $N_0$  atoms of  $A$  initially, plot the graph showing the variation of number of atoms of  $A$  and  $B$  versus time.

### K Thinking Process

Based on decay law of unstable radioactive nuclei.

**Ans.** Consider the situation shown in the graph.



At  $t = 0$ ,  $N_A = N_0$  (maximum) while  $N_B = 0$ . As time increases,  $N_A$  decreases exponentially and the number of atoms of  $B$  increases. They become  $(N_B)$  maximum and finally drop to zero exponentially by radioactive decay law.

**Q. 18** A piece of wood from the ruins of an ancient building was found to have a  $^{14}\text{C}$  activity of 12 disintegrations per minute per gram of its carbon content. The  $^{14}\text{C}$  activity of the living wood is 16 disintegrations per minute per gram. How long ago did the tree, from which the wooden sample came, die? Given half-life of  $^{14}\text{C}$  is 5760 yr.

### K Thinking Process

Carbon dating is a technique that uses the decay of carbon  $-14(^{14}\text{C})$  to estimate the age of organic materials, such as wood and leather.

**Ans.** Given,  $R = 12$  dis/min per g,  $R_0 = 16$  dis/min per g,  $T_{1/2} = 5760$  yr

Let  $t$  be the span of the tree.

According to radioactive decay law,

$$R = R_0 e^{-\lambda t} \quad \text{or} \quad \frac{R}{R_0} = e^{-\lambda t} \quad \text{or} \quad e^{\lambda t} = \frac{R_0}{R}$$

Taking log on both the sides

$$\lambda t \log_e e = \log_e \frac{R_0}{R} \Rightarrow \lambda t = \left( \log_{10} \frac{16}{12} \right) \times 2.303$$

$$t = \frac{2.303(\log 4 - \log 3)}{\lambda}$$

$$= \frac{2.303(0.6020 - 0.4771) \times 5760}{0.6931} \quad \left( \because \lambda = \frac{0.6931}{T_{1/2}} \right)$$

$$= 2391 \cdot 20 \text{ yr}$$

- Q. 19** Are the nucleons fundamental particles, or do they consist of still smaller parts? One way to find out is to probe a nucleon just as Rutherford probed an atom. What should be the kinetic energy of an electron for it to be able to probe a nucleon? Assume the diameter of a nucleon to be approximately  $10^{-15}$  m.

**K Thinking Process**

We have to use de-Broglie formula ( $\lambda = h / p$ ) to find momentum of the particle.

- Ans.** Each particle (neutron and proton) present inside the nucleus is called a nucleon.

Let  $\lambda$  be the wavelength  $\lambda = 10^{-15}$  m

To detect separate parts inside a nucleon, the electron must have wavelength less than  $10^{-15}$  m.

We know that

$$\lambda = \frac{h}{p} \text{ and KE} = \text{PE} \quad \dots(i)$$

$$\text{Energy} = \frac{hc}{\lambda} \quad \dots(ii)$$

From Eq. (i) and Eq. (ii),

$$\text{kinetic energy of electron} = \text{PE} = \frac{hc}{\lambda} = \frac{6.6 \times 10^{-34} \times 3 \times 10^8}{10^{-15} \times 1.6 \times 10^{-19}} \text{ eV}$$

$$\text{KE} = 10^9 \text{ eV}$$

- Q. 20** A nuclide 1 is said to be the mirror isobar of nuclide 2 if  $Z_1 = N_2$  and  $Z_2 = N_1$ . (a) What nuclide is a mirror isobar of  ${}_{11}^{23}\text{Na}$ ? (b) Which nuclide out of the two mirror isobars have greater binding energy and why?

**K Thinking Process**

Based on the mirror isobar concept and binding energy concept.

- Ans. (a)** According to question, a nuclide 1 is said to be mirror isobar of nuclide 2, if  $Z_1 = N_2$  and  $Z_2 = N_1$ .

Now in  ${}_{11}\text{Na}^{23}$ ,  $Z_1 = 11, N_1 = 23 - 11 = 12$

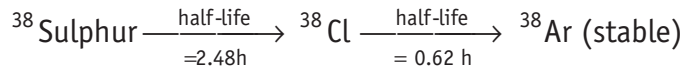
$\therefore$  Mirror isobar of  ${}_{11}\text{Na}^{23}$  is  ${}_{12}\text{Mg}^{23}$ , for which  $Z_2 = 12 = N_1$  and  $N_2 = 23 - 12 = 11 = Z_1$

- (b)** As  ${}_{12}^{23}\text{Mg}$  contains even number of protons (12) against  ${}_{11}^{23}\text{Na}$  which has odd number of protons (11), therefore  ${}_{11}^{23}\text{Na}$  has greater binding energy than  ${}_{11}\text{Na}^{23}$ .



## Long Answer Type Questions

**Q. 21** Sometimes a radioactive nucleus decays into a nucleus which itself is radioactive. An example is



Assume that we start with 1000  ${}^{38}\text{S}$  nuclei at time  $t = 0$ . The number of  ${}^{38}\text{Cl}$  is of count zero at  $t = 0$  and will again be zero at  $t = \infty$ . At what value of  $t$ , would the number of counts be a maximum?

### K Thinking Process

To solve this problem concept of chain of two decays will be used. For the process,

$A \rightarrow B \rightarrow C$ , the law of decay is  $\frac{dN_B}{dt} = \lambda_B N_B + \lambda_A N_A$ .

**Ans.** Consider the chain of two decays



At time  $t$ , Let  ${}^{38}\text{S}$  have  $N_1(t)$  active nuclei and  ${}^{38}\text{Cl}$  have  $N_2(t)$  active nuclei.

$$\frac{dN_1}{dt} = -\lambda_1 N_1 = \text{rate of formation of } \text{Cl}^{38}.$$

Also, 
$$\frac{dN_2}{dt} = -\lambda_2 N_2 + \lambda_1 N_1$$

But 
$$N_1 = N_0 e^{-\lambda_1 t}$$

$$\frac{dN_2}{dt} = \lambda_1 N_0 e^{-\lambda_1 t} - \lambda_2 N_2 \quad \dots(i)$$

Multiplying by  $e^{\lambda_2 t} dt$  and rearranging

$$e^{\lambda_2 t} dN_2 + \lambda_2 N_2 e^{\lambda_2 t} dt = \lambda_1 N_0 e^{(\lambda_2 - \lambda_1)t} dt$$

Integrating both sides

$$N_2 e^{\lambda_2 t} = \frac{N_0 \lambda_1}{\lambda_2 - \lambda_1} e^{(\lambda_2 - \lambda_1)t} + C$$

Since, at  $t = 0, N_2 = 0, C = -\frac{N_0 \lambda_1}{\lambda_2 - \lambda_1}$

$$\therefore N_2 e^{\lambda_2 t} = \frac{N_0 \lambda_1}{\lambda_2 - \lambda_1} (e^{(\lambda_2 - \lambda_1)t} - 1) \quad \dots(ii)$$

$$N_2 = \frac{N_0 \lambda_1}{\lambda_2 - \lambda_1} (e^{-\lambda_1 t} - e^{-\lambda_2 t})$$

For maximum count, 
$$\frac{dN_2}{dt} = 0$$

$$\lambda_1 N_0 e^{-\lambda_1 t} - \lambda_2 N_2 = 0 \quad \text{[From Eq. (i)]}$$

$$\Rightarrow \frac{N_0}{N_2} = \frac{\lambda_2}{\lambda_1} e^{\lambda_1 t} \quad \text{[From Eq. (ii)]}$$

$$e^{\lambda_2 t} - \frac{\lambda_2}{\lambda_1} \cdot \frac{\lambda_1}{(\lambda_2 - \lambda_1)} e^{\lambda_1 t} [e^{(\lambda_2 - \lambda_1)t} - 1] = 0$$

or 
$$e^{\lambda_2 t} - \frac{\lambda_2}{(\lambda_2 - \lambda_1)} e^{\lambda_2 t} + \frac{\lambda_2}{(\lambda_2 - \lambda_1)} e^{\lambda_1 t} = 0$$

$$\begin{aligned}
1 - \frac{\lambda_2}{(\lambda_2 - \lambda_1)} + \frac{\lambda_2}{(\lambda_2 - \lambda_1)} e^{(\lambda_1 - \lambda_2)t} &= 0 \\
\frac{\lambda_2}{(\lambda_2 - \lambda_1)} e^{(\lambda_1 - \lambda_2)t} &= \frac{\lambda_2}{(\lambda_2 - \lambda_1)} - 1 \\
e^{(\lambda_1 - \lambda_2)t} &= \frac{\lambda_1}{\lambda_2} \\
t &= \left( \log_e \frac{\lambda_1}{\lambda_2} \right) / (\lambda_1 - \lambda_2) \\
&= \frac{\log_e \left( \frac{2.48}{0.62} \right)}{2.48 - 0.62} \\
&= \frac{\log_e 4}{1.86} = \frac{2.303 \times 2 \times 0.3010}{1.86} \quad \left( \because \lambda = \frac{0.693}{T_{1/2}} \right) \\
&= 0.745 \text{ s}
\end{aligned}$$

**Note** Do not apply directly the formula of radioactive. Apply formulae related to chain decay.

- Q. 22** Deuteron is a bound state of a neutron and a proton with a binding energy  $B = 2.2$  MeV. A  $\gamma$ -ray of energy  $E$  is aimed at a deuteron nucleus to try to break it into a (neutron + proton) such that the  $n$  and  $p$  move in the direction of the incident  $\gamma$ -ray. If  $E = B$ , show that this cannot happen. Hence, calculate how much bigger than  $B$  must be  $E$  be for such a process to happen.

#### K Thinking Process

Apply conservation of energy as well as conservation of momentum.

- Ans.** Given binding energy  $B = 2.2$  MeV  
From the energy conservation law,

$$E - B = K_n + K_p = \frac{p_n^2}{2m} + \frac{p_p^2}{2m} \quad \dots(i)$$

From conservation of momentum,

$$p_n + p_p = \frac{E}{c} \quad \dots(ii)$$

As

$$E = B, \text{ Eq. (i) } p_n^2 + p_p^2 = 0$$

It only happen if  $p_n = p_p = 0$

So, the Eq. (ii) cannot satisfied and the process cannot take place.

Let  $E = B + X$ , where  $X \ll B$  for the process to take place.

Put value of  $p_n$  from Eq. (ii) in Eq. (i), we get

$$X = \frac{\left( \frac{E}{c} - p_p \right)^2}{2m} + \frac{p_p^2}{2m}$$

or 
$$2p_p^2 - \frac{2Ep_p}{c} + \frac{E^2}{c^2} - 2mX = 0$$

Using the formula of quadratic equation, we get

$$p_p = \frac{\frac{2E}{c} \pm \sqrt{\frac{4E^2}{c^2} - 8\left(\frac{E^2}{c^2} - 2mX\right)}}{4}$$

For the real value  $p_p$ , the discriminant is positive

$$\begin{aligned} \frac{4E^2}{c^2} &= 8\left(\frac{E^2}{c^2} - 2mX\right) \\ 16mX &= \frac{4E^2}{c^2} \\ X &= \frac{E^2}{4mc^2} \approx \frac{B^2}{4mc^2} \quad [ \because X \ll B \Rightarrow E \approx B ] \end{aligned}$$

- Q. 23** The deuteron is bound by nuclear forces just as H-atom is made up of  $p$  and  $e$  bound by electrostatic forces. If we consider the force between neutron and proton in deuteron as given in the form a coulomb potential but with an effective charge  $e'$

$$F = \frac{1}{4\pi\epsilon_0} \cdot \frac{e'^2}{r}$$

estimate the value of  $(e'/e)$  given that the binding energy of a deuteron is 2.2 MeV.

**Ans.** The binding energy is H-atom

$$E = \frac{me^4}{\pi\epsilon_0^2 h^2} = 13.6\text{eV} \quad \dots(i)$$

If proton and neutron had charge  $e'$  each and were governed by the same electrostatic force, then in the above equation we would need to replace electronic mass  $m$  by the reduced mass  $m'$  of proton-neutron and the electronic charge  $e$  by  $e'$ .

$$\begin{aligned} m' &= \frac{M \times N}{M + N} = \frac{M}{2} \\ &= \frac{1836m}{2} = 918m \end{aligned}$$

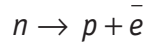
Here,  $M$  represents mass of a neutron/proton

$$\therefore \text{Binding energy} = \frac{918m (e')^4}{8\epsilon_0^2 h^2} = 2.2 \text{ MeV} \quad \dots(ii)$$

Dividing Eqs. (ii) and (i), we get

$$\begin{aligned} 918 \left(\frac{e'}{e}\right)^4 &= \frac{2.2 \text{ MeV}}{13.6 \text{ eV}} = \frac{2.2 \times 10^6}{13.6} \\ \left(\frac{e'}{e}\right)^4 &= \frac{2.2 \times 10^6}{13.6 \times 918} = 176.21 \\ \frac{e'}{e} &= (176.21)^{1/4} = 3.64. \end{aligned}$$

**Q. 24** Before the neutrino hypothesis, the beta decay process was thought to be the transition.



If this was true, show that if the neutron was at rest, the proton and electron would emerge with fixed energies and calculate them. Experimentally, the electron energy was found to have a large range.

**Ans.** Before  $\beta$ -decay, neutron is at rest. Hence,  $E_n = m_n c^2$ ,  $p_n = 0$

Or 
$$\mathbf{p}_n = \mathbf{p}_p + \mathbf{p}_e$$
  

$$\mathbf{p}_p + \mathbf{p}_e = 0 \Rightarrow |\mathbf{p}_p| = |\mathbf{p}_e| = p$$

Also,

$$E_p = (m_p^2 c^4 + p^2 c^2)^{\frac{1}{2}},$$

$$E_e = (m_e^2 c^4 + p^2 c^2)^{\frac{1}{2}}$$

$$= (m_e^2 c^4 + p_e^2 c^2)^{\frac{1}{2}}$$

From conservation of energy,

$$(m_p^2 c^4 + p^2 c^2)^{\frac{1}{2}} + (m_e^2 c^4 + p^2 c^2)^{\frac{1}{2}} = m_n c^2$$

$$m_p c^2 \approx 936 \text{ MeV}, m_n c^2 \approx 938 \text{ MeV}, m_e c^2 = 0.51 \text{ MeV}$$

Since, the energy difference between  $n$  and  $p$  is small,  $pc$  will be small,  $pc \ll m_p c^2$ , while  $pc$  may be greater than  $m_e c^2$

$$\Rightarrow m_p c^2 + \frac{p^2 c^2}{2m_p c^2} \approx m_n c^2 - pc$$

To first order  $pc \approx m_n c^2 - m_p c^2 = 938 \text{ MeV} - 936 \text{ MeV} = 2 \text{ MeV}$

This gives the momentum of proton or neutron. Then,

$$E_p = (m_p^2 c^4 + p^2 c^2)^{\frac{1}{2}} = \sqrt{936^2 + 2^2}$$

$$\approx 936 \text{ MeV}$$

$$E_e = (m_e^2 c^4 + p^2 c^2)^{\frac{1}{2}} = \sqrt{(0.51)^2 + 2^2}$$

$$\approx 2.06 \text{ MeV}$$

## Long Answer Type Questions

**Q. 25** The activity  $R$  of an unknown radioactive nuclide is measured at hourly intervals. The result found are tabulated as follows

$t(\text{h})$	0	1	2	3	4
$R(\text{MBq})$	100	35.36	12.51	4.42	1.56

- (i) Plot the graph of  $R$  versus  $t$  and calculate half-life from the graph.  
 (ii) Plot the graph of  $\ln\left(\frac{R}{R_0}\right)$  versus  $t$  and obtain the value of half-life from the graph.

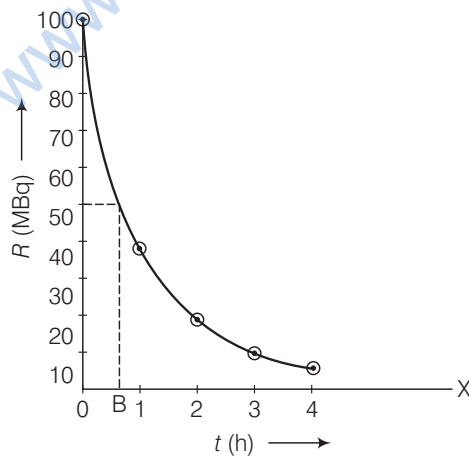
### K Thinking Process

Based on Decay law and half-life.

**Ans.** In the table given below, we have listed values of  $R(\text{MBq})$  and  $\ln\left(\frac{R}{R_0}\right)$ .

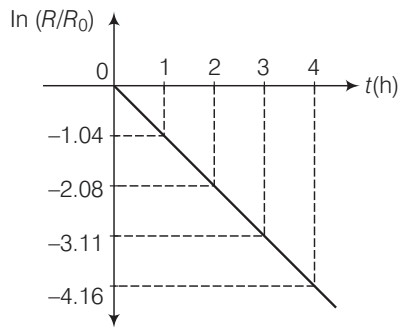
$t(\text{h})$	0	1	2	3	4
$R(\text{MBq})$	100	35.36	12.51	4.42	1.56
$\frac{R}{R_0}$	—	-1.04	-2.08	-3.11	-4.16

- (i) When we plot the graph of  $R$  versus  $t$ , we obtain an exponential curve as shown.



From the graph we can say that activity  $R$  reduces to 50% in  $t = OB \approx 40$  min  
 So,  $t_{1/2} \approx 40$  min.

(ii) The adjacent figure shows the graph of  $\ln(R/R_0)$  versus  $t$ .



Slope of this graph  $= -\lambda$

from the graph,

$$\lambda = -\left(\frac{-4.16 - 3.11}{1}\right) \Rightarrow = 1.05\text{h}^{-1}$$

Half-life

$$T_{1/2} = \frac{0.693}{\lambda} = \frac{0.693}{1.05} = 0.66\text{h}$$

$$= 39.6 \text{ min} \approx 40 \text{ min}$$

**Q. 26** Nuclei with magic number of proton  $Z = 2, 8, 20, 28, 50, 52$  and magic number of neutrons  $N = 2, 8, 20, 28, 50, 82$  and  $126$  are found to be very stable.

(i) Verify this by calculating the proton separation energy  $S_p$  for  $^{120}\text{Sn}(Z=50)$  and  $^{121}\text{Sb}(Z=51)$ .

The proton separation energy for a nuclide is the minimum energy required to separate the least tightly bound proton from a nucleus of that nuclide. It is given by

$$S_p = (M_{Z-1, N} + M_H - M_{Z, N})c^2.$$

Given,  $^{119}\text{In} = 118.9058 \text{ u}$ ,  $^{120}\text{Sn} = 119.902199 \text{ u}$ ,  
 $^{121}\text{Sb} = 120.903824 \text{ u}$ ,  $^1\text{H} = 1.0078252 \text{ u}$ .

(ii) What does the existence of magic number indicate?

**Ans. (i)** The proton separation energy is given by

$$S_{p\text{Sn}} = (M_{119,70} + M_H - M_{120,70})c^2$$

$$= (118.9058 + 1.0078252 - 119.902199)c^2$$

$$= 0.0114362c^2$$

Similarly

$$S_{p\text{Sb}} = (M_{120,70} + M_H - M_{121,70})c^2$$

$$= (119.902199 + 1.0078252 - 120.903822)c^2$$

$$= 0.0059912c^2$$

Since,  $S_{p\text{Sn}} > S_{p\text{Sb}}$ , Sn nucleus is more stable than Sb nucleus.

(ii) The existence of magic numbers indicates that the shell structure of nucleus similar to the shell structure of an atom. This also explains the peaks in binding energy/nucleon curve.